

## Nickel Sorption onto Chitosan - Silica Hybrid Aerogel from Aqueous Solution

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### Abstract

The utilization and efficiency of agricultural waste-derived chitosan-silica hybrid aerogel “(CS)<sub>h</sub>A” on nickel removal from aqueous solution was examined and optimum parameters for pH, contact time, and initial ion concentration were determined during batch sorption system studies. Metal recovery was performed on the adsorbent using separately dilute concentrations of hydrochloric acid, acetic acid, ammonia, and sodium hydroxide solutions as eluents. The results generated were analyzed from kinetic and isotherm studies. The maximum Ni<sup>2+</sup> adsorption of 99.78 % was established at 60 min and pH 3 in this study. The batch studies revealed that the percentage of nickel ion removal by the adsorbent decreased along with an increase in the initial Ni<sup>2+</sup> ions concentration. The pseudo-second-order, the best fit of the kinetic model, has the values of its correlation coefficient “R<sup>2</sup>” ranging from 0.9 to 1, whereas the Langmuir model which had the maximum monolayer adsorption capacity of 85.84 mg g<sup>-1</sup>, was the best isotherm in interpreting the sorption process and the calculated separation factor was higher than 0 but less than 1. Dilute hydrochloric acid (0.1 M) was the best eluent in removing bound nickel ions (55.63 %) from (CS)<sub>h</sub>A.

**Keywords:** Adsorbent, Chitosan - silica hybrid aerogel, Isotherm, Kinetic, Nickel

### Introduction

Heavy metals exposure can bio-accumulate ending up permanently in the environment which could be a serious threat to humanity and the ecological system [1-3]. Metals like Pb<sup>2+</sup>, Hg<sup>2+</sup>, Cd<sup>2+</sup>, Ni<sup>2+</sup>, and Cr<sup>6+</sup> can be highly toxic and their toxicity level is dependent on the metal type, valency, biological role, volume, and concentration [4]. Heavy metal ions released from nickel-related activities contaminate water [5,6]. This is a universal problem that is receiving attention globally [7] because of its hazardous effect on soil and water [4,8].

According to WHO [9], nickel is of concern and its concentration must be reduced to a permissible level before released to the environment [7]; a concentration of nickel II above 0.02 mg/L can result in adverse health effects, including nausea, vomiting, diarrhea, pulmonary fibrosis, renal edema, and skin dermatitis [10,11]. It is, therefore, important to treat Ni (II) rich effluent before releasing it to the environment. Heavy metals toxicity has received high sensitization and hence research into technologies

targeted at water cleanup. These technologies, like chemical precipitation, adsorption, reverse osmosis, ultra filtration, or ion exchange, have been investigated and some found to be expensive and generating harmful products [10,12]. However, adsorption proves to be less expensive, highly efficient and there is the ease in its design and usage [13-15]. Other merits of this technique involve low residue generation, easy metal recovery, and probability for reusability of sorbent [16]. Functional groups, e.g., carbonyl, amino, hydroxyl, phosphate, and sulfate existing in the adsorbent, can enhance the sorption of heavy metals likewise also the adsorption of sorbate unto the surface, pores, and complexation by the adsorbent [17].

Gottipati and Mishra [18] reported that surface area and pore volume are determining factors for sorption efficiency. Also, Gong *et al.* [19] reported that high surface area leads to metal ions adhering to the adsorbent. Biopolymer-based aerogels have continued to receive attention especially chitosan-based aerogels [20-22]. The presence of amine and hydroxyl groups in aerogels makes them highly effective for adhering to both anionic pollutants and heavy metal ions [23].

Besides, sorption capacity can be optimized by controlling their porosity and surface area [24]. Ebisike *et al.* [25] produced a high surface area chitosan-silica hybrid aerogel which possesses hydroxyl, carbonyl, amino, siloxane, and silanol groups which can boost the sorption of metal ions. Adsorption of cadmium using chitosan-silica hybrid aerogel had been reported as successful [26] but there has not been any report on the application of this aerogel for adsorption of nickel, which is a highly prevalent heavy metal in wastewaters in the study environment.

This work will provide information on the utilization of chitosan-silica hybrid aerogel in nickel ions adsorption. To obtain the optimum sorption conditions, the adsorption process will be analyzed on pH, contact time, and initial nickel ion concentrations. Similarly, the desorption of the nickel ions bound by the adsorbent. Kinetic and sorption models will also be determined.

## Materials and methods

### Adsorbent preparation

Chitosan-silica hybrid aerogel [(CS)<sub>h</sub>A] was prepared following the method by Ebisike *et al.* [25]. Adsorption studies were carried out on nickel ion from an aqueous solution using the derived aerogel.

### pH profile studies

Hydrochloric acid and sodium hydroxide were used to adjust the solutions to their respective labeled pH 1 to 8. The adsorbent was weighed into the 8 containers containing the solutions and then agitated for 6 h, after which the resultant supernatant was collected for analysis.

### Contact time studies

The (CS)<sub>h</sub>A was constantly agitated with 10 ml Ni<sup>2+</sup> solution of different concentrations (that is, 10, 20, 50, 75, 100 and 200 mg/L). These solutions had been regulated to optimum pH. Thereafter, 2 aliquots (2 mL) were removed from each suspension and transferred into uncontaminated beakers for studies at 5, 10, 15, 30, 60, 120, 180, 240, 300, and 360 min. The supernatants were gotten after centrifugation and then decanted into uncontaminated sample bottles for analysis.

### Metal recovery studies

The (CS)<sub>h</sub>A was subjected to desorption studies using eluents of various concentrations (0.1, 0.05, and 0.01 M) of hydrochloric acid, acetic acid, ammonia, and sodium hydroxide solutions, respectively. Sorbent material (0.02 g) was weighed and put in different beakers containing 10 mL of 50 mg/L Ni<sup>2+</sup> solutions which had already been regulated to pH 3. This was stirred continuously for the optimum sorption time duration. The sorbent was centrifuged and the residue was introduced separately into 10 ml of 0.01, 0.05, and 0.1 M respectively for HCl, CH<sub>3</sub>COOH, NH<sub>3</sub>, and NaOH. This was stirred up till optimum time and the supernatant was collected for analysis.

## Results and discussion

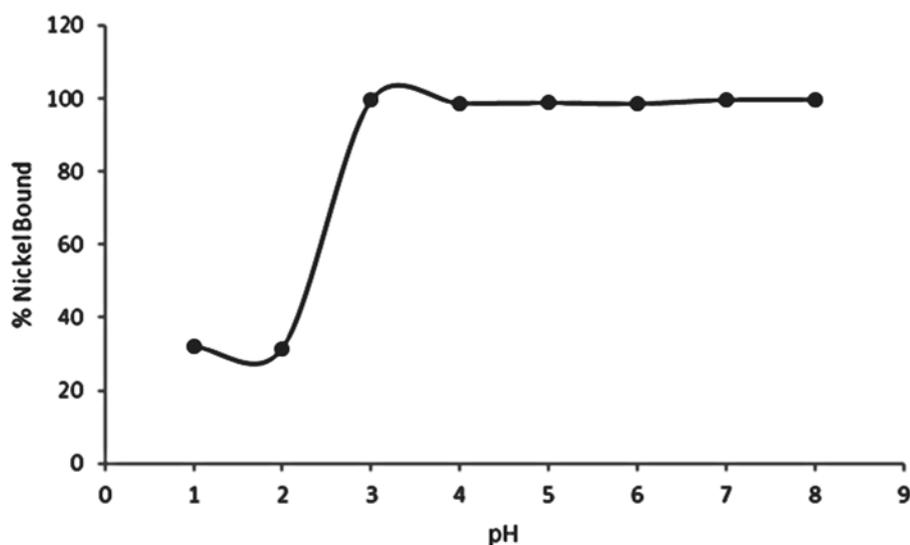
### Effect of pH

**Figure 1** shows the percentage of nickel ion bound by chitosan-silica hybrid aerogel. An optimum % Ni<sup>2+</sup> bound (99.89 %) was reached at pH 3 with reduced nickel ion bound observed below this pH which could be as a result of more hydrogen ions surrounding the binding sites of the derived aerogel, hence making sorption to be unfavorable. Givianrad *et al.* [27] report that the more the H<sup>+</sup> ions on binding sites of the sorbent at lower pH, the more the desorption of bound metals from the sites. It has been reported that when the ionic state of the amino and carboxyl groups of a sorbent is in a strongly acidic medium (pH < 3), it will be protonated but becomes deprotonated at pH ≥ 3 [28-30]. Ebisike *et al.* [25] reported amino groups as part of the functional groups in the derived aerogel.

### Effect of initial concentration

Sorption of initial nickel concentrations (10 - 200 mg/L) onto (CS)<sub>n</sub>A was carried out at constant contact time, temperature, and optimum pH and is presented in **Figure 2**. The % Ni<sup>2+</sup> ion removed is low at high initial concentration but increases when initial concentration reduces. Ebisike *et al.* [26] reported a similar trend with initial concentration studies during the adsorption studies of Cd<sup>2+</sup> unto chitosan-silica hybrid aerogel. This increase in nickel ions removed after the initial ion concentration reduces may be linked to the formation of a mono ionic layer on the sorbent surface which may hinder the formation of an extra sorbate layer on the sorbent [31].

In addition, the adsorption sites ideally become saturated as the number of nickel ions removed from solution increases with increasing initial nickel concentration until a plateau-like shape is formed indicating that further nickel ions adsorption cannot occur [32]. The maximum concentration of Ni<sup>2+</sup> removed at 10, 20, 50, 75, 100, and 200 mg/L are 4.98, 9.96, 24.95, 37.26, 48.70, and 85.84 mg/g respectively. These values are as presented in **Table 1**.



**Figure 1** Plot of percentage nickel bound versus pH

### Effect of contact time

Figure 3 shows the results of 50 mg/L percent of nickel that is held on the surface of (CS)<sub>h</sub>A with respect to time. It was observed that 73.52 % Ni<sup>2+</sup> was bounded by the adsorbent at the 5<sup>th</sup> min of contact. This implies a 14.70 %min<sup>-1</sup> sorption rate. Such a sorption rate points to chemisorption as the possible reaction mechanism [16,33]. Also from the plot there looked to be adsorption and desorption of nickel ions on the chitosan-silica hybrid aerogel. The % nickel removal rises as contact time increases until equilibrium is attained. Sixty min was achieved as the optimum sorption time for this study.

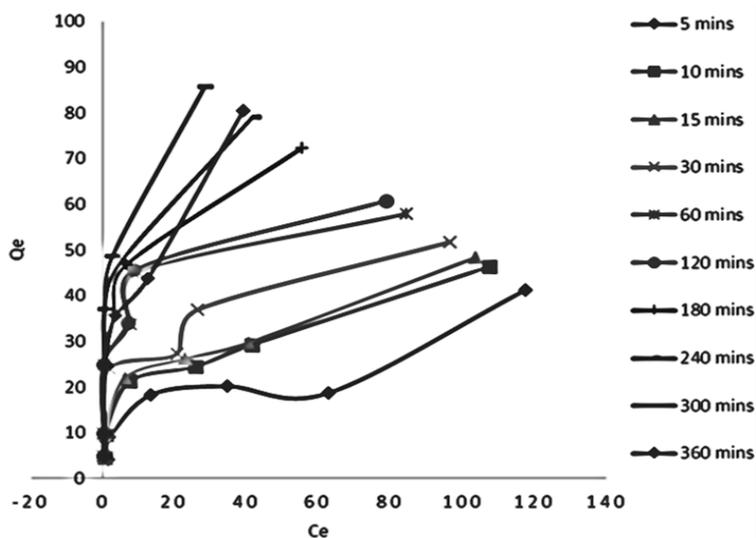


Figure 2 Plot of the amount of nickel adsorbed as a function of various concentration over a period of time.

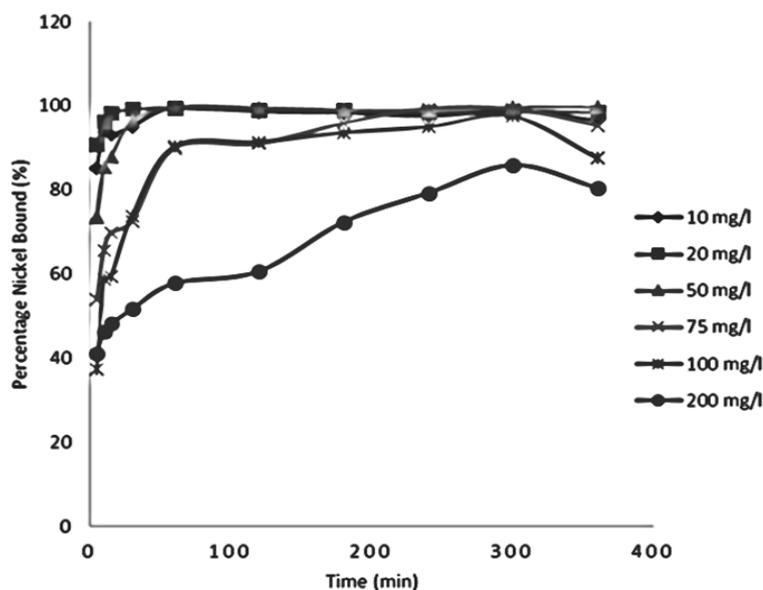


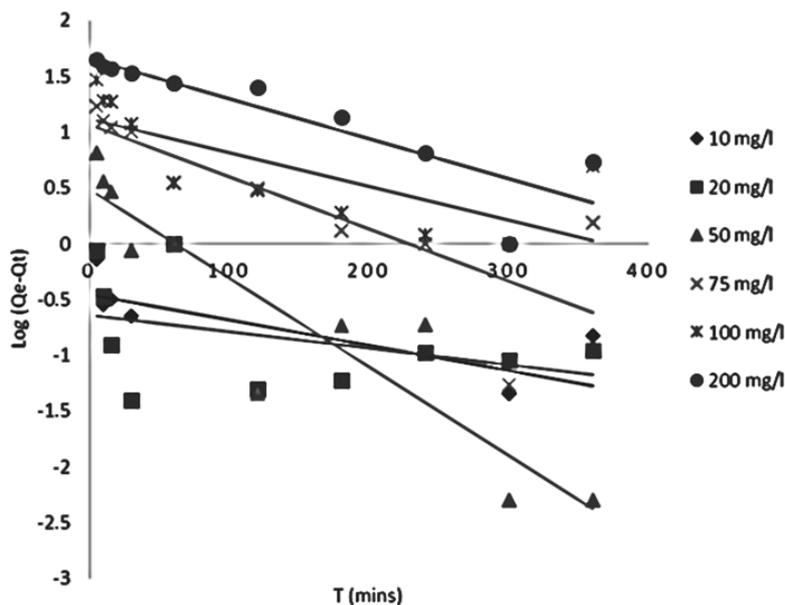
Figure 3 Plot of the percentage of various nickel bound concentration study as a function of time.

**Table 1** Kinetic parameters for nickel sorption onto chitosan - silica hybrid aerogel.

Parameters	10 ppm	20 ppm	50 ppm	75 ppm	100 ppm	200 ppm
<b>Pseudo-first-order</b>						
q <sub>e.exp</sub> (mg g <sup>-1</sup> )	4.9750	9.9550	24.9450	37.2550	48.7000	85.8400
q <sub>e.cal</sub> (mg g <sup>-1</sup> )	0.3525	0.2311	3.0733	11.6332	13.1917	46.0045
k <sub>1</sub> (min <sup>-1</sup> )	0.0053	0.0035	0.0182	0.0108	0.5595	0.0083
R <sup>2</sup>	0.3964	0.1557	0.8488	0.6771	0.0069	0.7940
<b>Pseudo-second-order</b>						
q <sub>e.exp</sub> (mg g <sup>-1</sup> )	4.9750	9.9550	24.9450	37.2550	48.7000	85.8400
q <sub>e.cal</sub> (mg g <sup>-1</sup> )	4.8733	9.8522	25.0000	37.1747	46.9484	84.7458
k <sub>2</sub> (g mg <sup>-1</sup> min <sup>-1</sup> )	-1.8228	-0.8959	0.0279	0.0042	0.0044	0.0006
R <sup>2</sup>	0.9994	1.0000	1.0000	0.9987	0.9947	0.9877
<b>Intra-particle diffusion</b>						
k <sub>id</sub> (mg g <sup>-1</sup> min <sup>-1/2</sup> )	0.0051	0.0004	0.2477	0.7958	1.4073	1.9679
C (mg g <sup>-1</sup> )	4.6339	9.5234	20.2480	23.3590	25.6090	30.5750
R <sup>2</sup>	0.7743	0.6273	0.8609	0.8909	0.7810	0.9786
<b>Elovich</b>						
A <sub>E</sub> (g mg <sup>-1</sup> )	8.9847	8.6505	0.7768	0.2578	0.1617	0.1002
B <sub>E</sub> (mg g <sup>-1</sup> min <sup>-1</sup> )	8.09E + 15	9.98E + 33	1778354	197.5908	57.2926	84.5465
R <sup>2</sup>	0.2496	0.4561	0.7833	0.9504	0.8751	0.9177

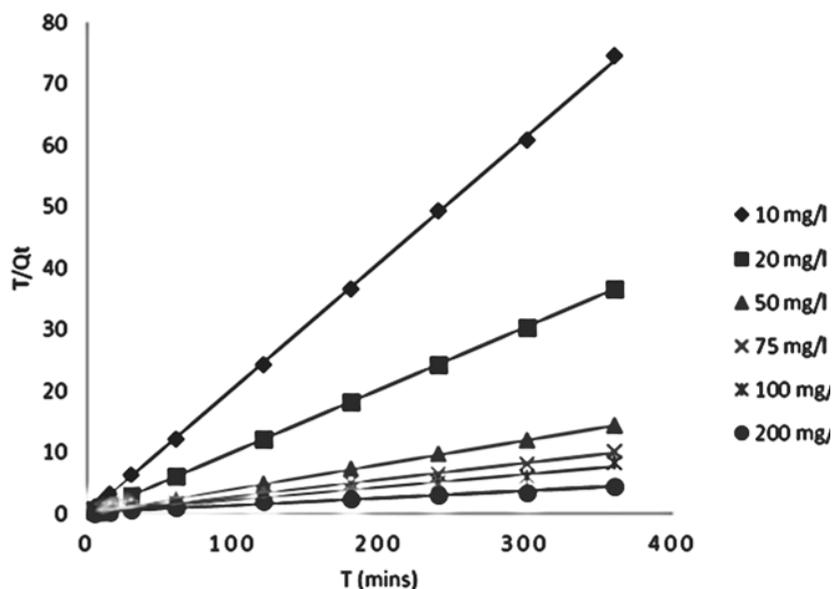
The sorption process was subjected to pseudo-first<sup>st</sup>-order or Lagergen kinetic, pseudo-second-order, intra-particle diffusion, and Elovich kinetic models. **Table 1** gives the different kinetic model parameters for different test conditions used in explaining the batch sorption process.

The applicability of the pseudo-first-order or Lagergen kinetic model is a result of the high value of the correlation coefficient (R<sup>2</sup>) obtained from the line plot. The correlation coefficient value for 50 and 200 mg/L is high whereas poor linearity is observed for other studied nickel concentrations which record low values as shown in **Figure 4**. Hence, there is a deviation of this sorption process from this model. Also, to support the applicability of this model, the value of the experimental equilibrium uptake must be equivalent to the calculated, but in this case, they are far apart as seen in **Table 1**, and therefore cannot be used to interpret the kinetics. Information for this model (**Figure 5**) and the derived parameter values (**Table 1**) shows that the R<sup>2</sup> value is above 0.9 and to some extent unity for some studied concentration. This means that this model can explain the adsorption process.



**Figure 4** Pseudo-first-order plot for derived chitosan-silica hybrid aerogel in a nickel solution.

Moreover, the acceptability of the high value of the  $R^2$  in explaining the sorption process is justified when there is a close relationship between the test ( $q_{e,exp}$ ) and calculated ( $q_{e,cal}$ ) equilibrium adsorptions. Also, with this model, the rate-determining step is controlled by chemisorption having forces that share or interchange electrons amongst  $Ni^{2+}$  and  $(CS)_nA$ . The values of  $k_2$  for 10 and 20 mg/L are negative which could imply that the rate of chemical sorption is reducing as time increases. Some literature (**Table 2**) provides information concerning the Pseudo-second-order kinetic model [34-36] for the adsorption of  $Ni^{2+}$  present in an aqueous solution.

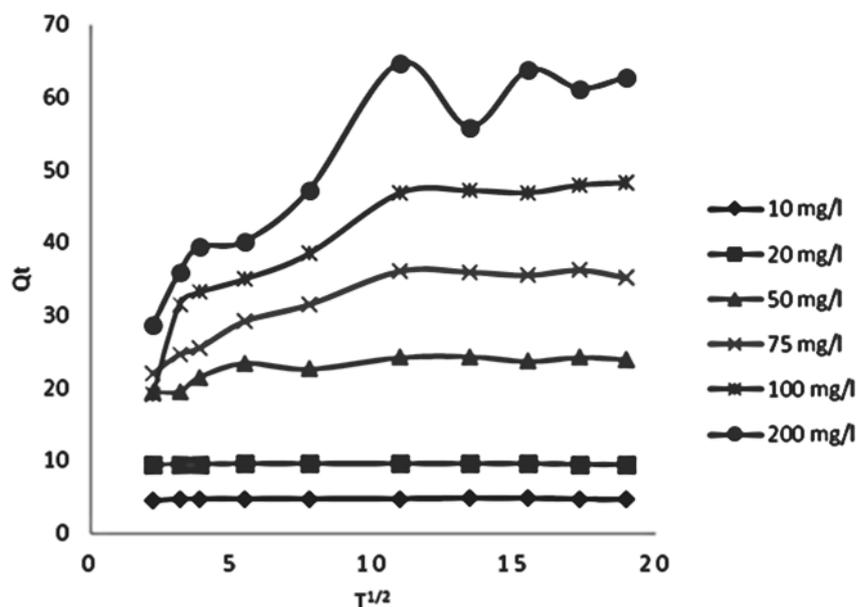


**Figure 5** Pseudo-second-order plot for derived chitosan - silica hybrid aerogel in nickel solution.

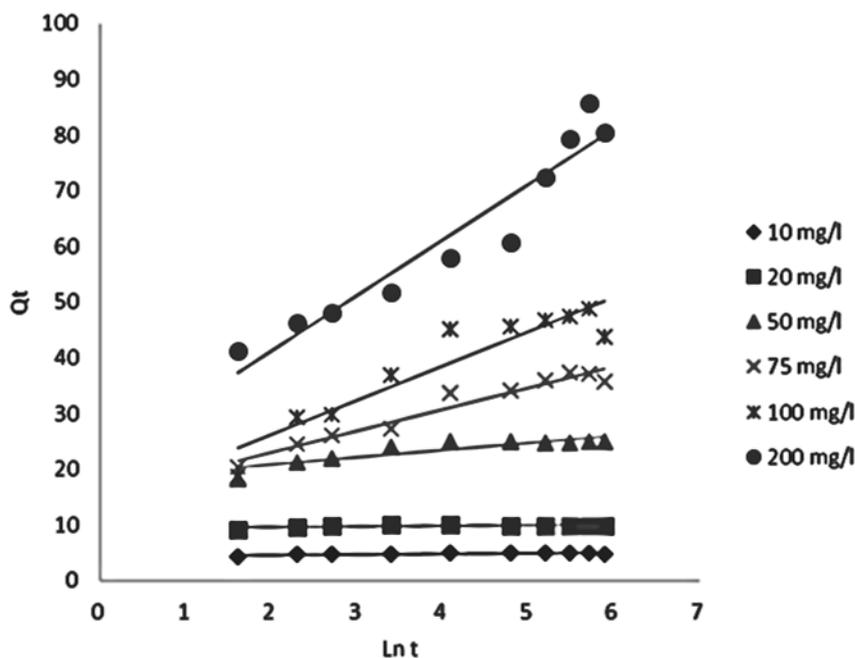
**Table 1** and **Figure 6** give the parameters for intra-particle diffusion and its plot respectively. Analyzing the plot, it was observed that 50, 75, 100, and 200 mg/L are not completely linear. There are 3 distinct linear parts to explain the sorption steps which are; instantaneous adsorption observed at the initial stage, intra-particle diffusion as the 2<sup>nd</sup> stage, whereas in the last stage there is a decline in intra-particle diffusion as a result of very low adsorbate concentration left in solution [37,38]. This last stage is what was observed for both the 10 and 20 mg/L sorption process. From the plot, the straight lines did not pass through the origin which means that intra-particle diffusion cannot be said to be the only rate-limiting step [37,39].

**Figure 7** is the Elovich plot for the sorption process. The correlation coefficients for 50, 75, 100, and 200 mg/L are greater than 0.8 which is also a good fit to the model and shows the significance of diffusional rate-limiting step in nickel sorption onto (CS)<sub>h</sub>A, even though the test data at low concentrations of 10 and 20 mg/L cannot explain this model. Also, an increase in the initial concentration of the solution from 10 to 200 mg/L causes the desorption constant to reduce.

The adsorption process which explains how metal ions concentration relates to adsorbents were also subjected to Langmuir, Freundlich, Temkin, and Dubinin-Raduskevich Isotherm. It can also improve the experimental conditions set to actualize optimum uptake of metal [40,41] as observed in the result reported.



**Figure 6** Intra-particle diffusion model plot for derived chitosan-silica hybrid aerogel in a nickel solution.



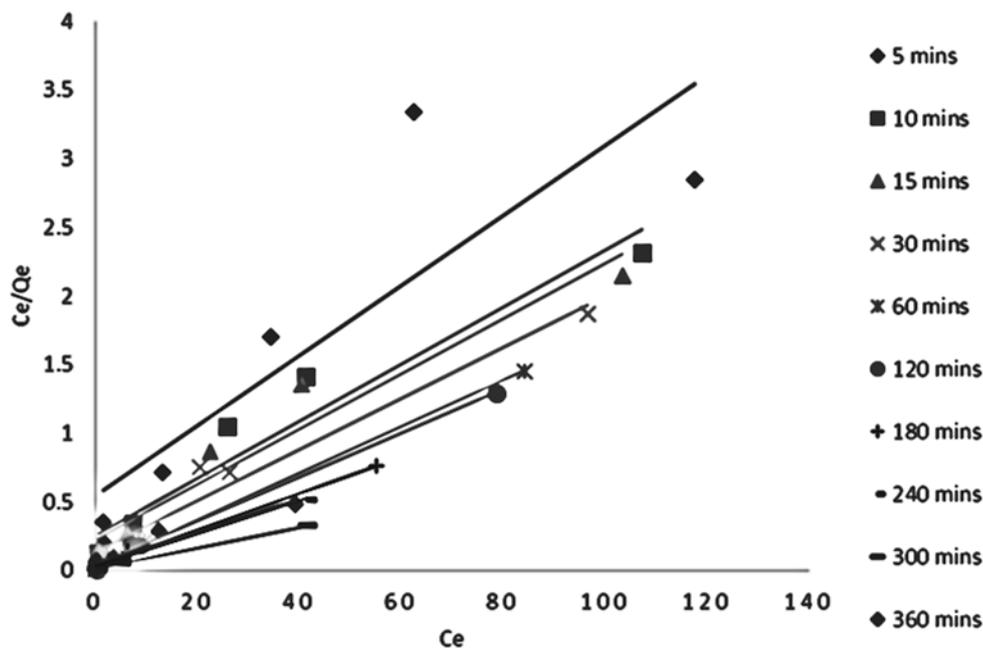
**Figure 7** Elovich model plot for derived chitosan - silica hybrid aerogel in a nickel solution.

**Table 2** Comparison for the removal of Ni<sup>2+</sup> by different adsorbents.

Adsorbent	Sorption capacity (mg/g)	Kinetic model	Isotherm model	Reference
Lignocellulose/montmorillonite (LNC/MMT) nanocomposite	94.86	Pseudo-second-order	Langmuir	Zhang and Wang, 2015
Activated Carbon Olive Stone	77.00	Pseudo-second-order	Freundlich	Ahmed <i>et al.</i> , 2019
Citric acid modified <i>Ceiba pentandra</i> hulls	34.34	Pseudo-second-order	Langmuir	Ramana <i>et al.</i> , 2012
Amine-functionalized silica aerogel	40.32	-	Langmuir	Sertsing <i>et al.</i> , 2018
Ulva lactuca biomass	38.28	Pseudo-second-order	Langmuir	Long <i>et al.</i> , 2018
Modified holly sawdust	22.47	Pseudo-second-order	Langmuir	Samarghandi, <i>et al.</i> , 2011
Tamarid bark	15.34	Pseudo-second-order	Langmuir	Zang <i>et al.</i> , 2010
Modified Riverbed Sand	0.86	Pseudo-first-order	Langmuir	Abdullah and Prasad, 2010
Mustard oil cake	2.761	Pseudo-second-order	Temkin	Khan <i>et al.</i> , 2012
	2.956	Pseudo-second-order	Temkin	
	3.208	Pseudo-second-order	Freundlich	

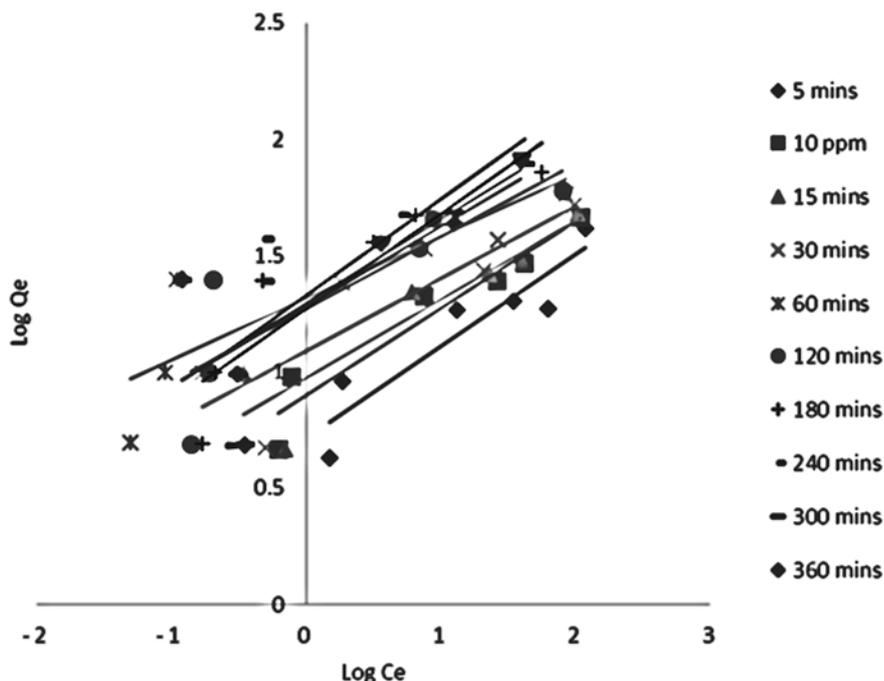
Adsorbent	Sorption capacity (mg/g)	Kinetic model	Isotherm model	Reference
Strains of <i>Yarrowia lipolytica</i> ;	95.33	Pseudo-second-order	Langmuir and Freundlich	Shinde <i>et al.</i> , 2012
NCIM 3589	85.44	Pseudo-second-order		
NCIM 3590				
Calcium Alginate/ H-magadiite-D2EHPA (CAM-D2EHPA)	44	Pseudo-second-order	Langmuir	Attar <i>et al.</i> , 2019
Peat	61.27	Pseudo-second-order	Langmuir	Przemysław <i>et al.</i> , 2015
<i>Callinectes sapidus</i> biomass	29.15	Pseudo-second-order	Freundlich	Foroutan <i>et al.</i> , 2019
Magnetic activated carbon/chitosan composite (MACCS) beads	108.70	Pseudo-second-order	Langmuir	Le <i>et al.</i> , 2019
Activated carbon from sewage sludge	11.52	Pseudo-second-order	Freundlich	Khelifi, <i>et al.</i> , 2018
Chitin/lignin hybrid material	70.41	Pseudo-second-order	Langmuir	Bartczak <i>et al.</i> , 2017
Chitosan-silica hybrid aerogel	85.84	Pseudo-second-order	Langmuir	Present study

**Figures 8 - 11** and **Table 3** are representations of investigations of sorption isotherms of various initial Ni<sup>2+</sup> concentrations. While most of the correlation coefficient values of the sorption isotherms are high, that of Langmuir (that is: 0.748 - 0.999) could be considered best in interpreting what is happening in the system. In other words, the sorption process is homogeneous [42] and every other adsorption stops after a single layer of ion had been formed on the surface of the adsorbent. The Dimensionless separation factor 'R<sub>L</sub>' was found within the range from 0.011 - 0.682 which is an indication of the applicability of Langmuir isotherm on the sorption of Ni<sup>2+</sup> onto this adsorbent. The R<sub>L</sub> values reduce as the initial nickel ion concentration increases, indicating more favorable adsorption at a higher Ni<sup>2+</sup> concentration. Maximum sorption can be considered for optimization purposes because it points out the best uptake capacity of adsorbent [43] The optimum nickel adsorption values 'Q<sub>max</sub>' obtained at initial concentration were recorded from 39.22 to 131.60 mg/g. **Table 2** shows a similar report from the literature supporting the Langmuir model for nickel ion sorption from aqueous solution [44-46].



**Figure 8** Langmuir isotherm plot for derived chitosan-silica hybrid aerogel in a nickel solution.

Freundlich isotherm has its correlation coefficient values in the range (0.596 - 0.919) of which 240, 300 and 360 min have their  $R^2$  values as 0.647, 0.619, and 0.596, respectively to be moderately correlated while other studied times are highly correlated. Also, a Freundlich constant 'n' which is a measure of the deviation of adsorption from linearity is above 1 but less than 10 implying also the application of Freundlich isotherm interpreting the sorption process. It was reported that adsorption intensity is more heterogeneous as it approaches 0 [47], while it is considered homogeneous if it tends to unity [16]. The slope  $1/n$  lies above zero but less than one and all values are closer to 0 as contact time increases, hence, it is more homogenous.



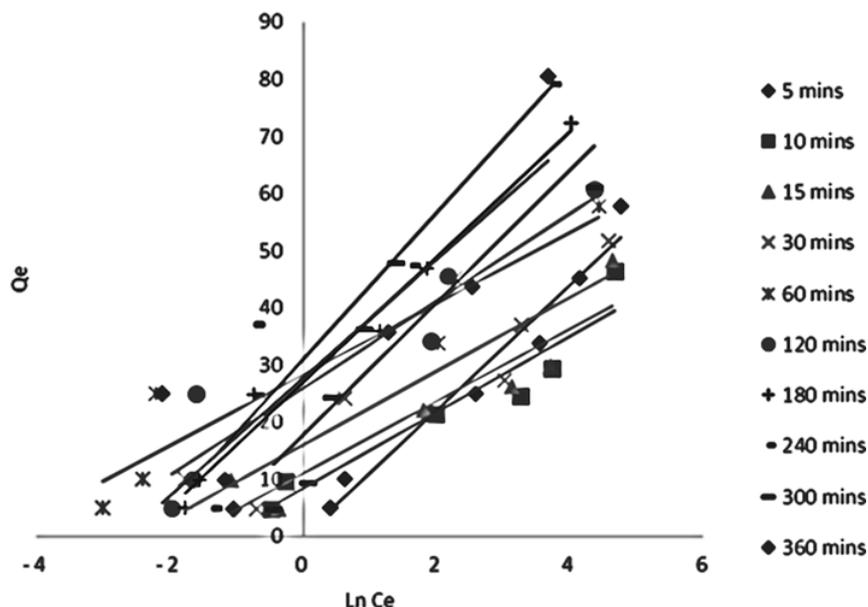
**Figure 9** Freundlich Isotherm plot for derived chitosan-silica hybrid aerogel in a nickel solution.

Foroutan *et al.* [48] when using *Callinectes sapidus* biomass to remove nickel ions from aqueous stream predicted Freundlich as the best model in interpreting the adsorption process. Similarly, Freundlich isotherm was also predicted to be the best fit when nickel ion was adsorbed by Activated carbon olive stone [49].

The  $R^2$  value of Temkin falls in the range of 0.867 to 0.987, indicating that they are highly correlated. This implies that the interaction between the (CS)<sub>h</sub>A and the Ni<sup>2+</sup> as well as the adsorption process relates to surface coverage [37].

Dubinini Raduskevick isotherm model with its  $R^2$  value falling in the range of 0.844 to 0.953 could interpret the adsorption process, unlike the lowly correlated values. The calculated mean free energy of adsorption 'E' for all the time frames studied is less than 8 kJ/mol (0.032 - 3.540 kJ/mol). This process relates to physisorption.

A comparison of different adsorbents with respect to sorption efficiency achieved for removal of nickel ions is presented in **Table 2**. In this study, chitosan-silica hybrid aerogel showed better sorption capacity for Ni<sup>2+</sup> as compared to other adsorbents. It should be noted that the sorption of nickel onto chitosan-silica hybrid aerogel is a function of the adsorbent, the radius of the spherical shell of the solvent molecule formed around the Ni<sup>2+</sup> ion, the equilibrium constant for hydrolysis reaction, electronegativity, and relative binding affinity [50,51]. Hence, these factors might competitively and effectively influence the sorption of other heavy metals onto the adsorbent having been successfully utilized for cadmium removal [26]. This is a sound basis for this research which aims at investigating the remover efficiency of Ni<sup>2+</sup> by chitosan-silica hybrid aerogel.



**Figure 10** Temkin isotherm plot for derived chitosan - silica hybrid aerogel in a nickel solution.

**Table 3** Isotherm parameters of a different model for nickel ions sorption onto chitosan-silica hybrid aerogel.

Parameters		Time (min) for Ni adsorption onto chitosan-silica hybrid Aerogel									
		5	10	15	30	60	120	180	240	300	360
<b>Langmuir</b>	$q_{max}$ (mg/g)	39.22	48.08	49.51	53.48	57.47	62.11	75.19	83.33	83.10	84.75
	$k_L$ (L/mg)	1.832	3.945	4.521	7.418	22.57	27.62	31.65	35.21	34.81	18.59
	$R_L$										
	10 mg/L	0.682	0.549	0.523	0.419	0.203	0.184	0.192	0.191	0.175	0.313
	20 mg/L	0.517	0.379	0.354	0.265	0.113	0.101	0.106	0.106	0.096	0.186
	50 mg/L	0.300	0.196	0.180	0.126	0.049	0.043	0.045	0.045	0.041	0.084
	75 mg/L	0.222	0.140	0.127	0.088	0.033	0.029	0.031	0.031	0.028	0.057
	100 mg/L	0.176	0.109	0.099	0.067	0.025	0.022	0.023	0.023	0.021	0.044
	200 mg/L	0.097	0.057	0.052	0.035	0.013	0.011	0.012	0.012	0.011	0.022
$R^2$	0.748	0.937	0.937	0.923	0.999	0.996	0.997	0.993	0.995	0.934	
<b>Freundlich</b>	$k_F$ (mg/g)L/ mg	5.202	7.989	9.395	12.32	20.79	18.66	18.79	21.35	19.86	19.24
	$n$	2.535	2.677	2.948	3.195	3.755	3.198	2.447	2.416	2.777	2.901
	$L/n$	0.395	0.374	0.339	0.313	0.266	0.313	0.409	0.414	0.360	0.345
	$R^2$	0.853	0.919	0.84	0.770	0.764	0.734	0.833	0.647	0.619	0.596
<b>Temkin</b>	$B$ (mg/L)	10.93	6.636	6.320	6.450	6.215	7.642	10.99	12.86	11.63	10.36
	$b_T$	229.1	377.1	396.0	388.0	402.7	3.398	227.7	194.6	215.3	241.6
	$A_T$ (L/mg)	1.021	3.539	5.697	12.07	94.54	327.5	11.72	11.14	4.566	14.72
	$R^2$	0.987	0.974	0.980	0.867	0.902	0.899	0.982	0.900	0.813	0.763
<b>Dubinin-Radushkevich</b>	$q_{max}$ (mg/g)	23.71	33.34	25.88	27.42	47.23	48.25	51.39	72.97	51.75	49.49
	$k_{ad}$ (mg/kj <sup>2</sup> )	-1E-6	-2E-7	-1E-7	-6E-8	-4E-8	-7E-8	-9E-8	-2E-7	-5.00	-6E-7
	$E$ (Kj/mol)	0.707	1.58	2.24	2.89	3.54	2.67	2.36	1.58	0.032	0.913
	$R^2$	0.843	0.728	0.482	0.413	0.884	0.816	0.947	0.900	0.953	0.938

**Metal recovery studies**

Figure 12 is a plot showing the effect of various concentrations (0.1, 0.05, and 0.01 M) of eluents on the recovery of Ni<sup>2+</sup> present on chitosan-silica hybrid aerogel respectively. HCl (0.1 M) was most effective in desorbing Ni<sup>2+</sup> from (CS)<sub>h</sub>A amongst other eluents that were applied for the study. This is because the maximum removal of bound Ni<sup>2+</sup> from this adsorbent by HCl, CH<sub>3</sub>COOH, NH<sub>3</sub>, and NaOH was recorded as 55.63, 34.72, 1.21, and 1.14 % respectively at 60 min. The acidic eluents cause the derived aerogel surface to be protonated, hence, the nickel bound ions will be desorbed [52,53] whereas the alkaline eluents had a very low effect on the desorption of nickel ions. This low effect may be attributed to more hydroxyl groups formed in the solution making more openings at the active sites which support electrostatic exchanges amongst cations and sorbent surface.

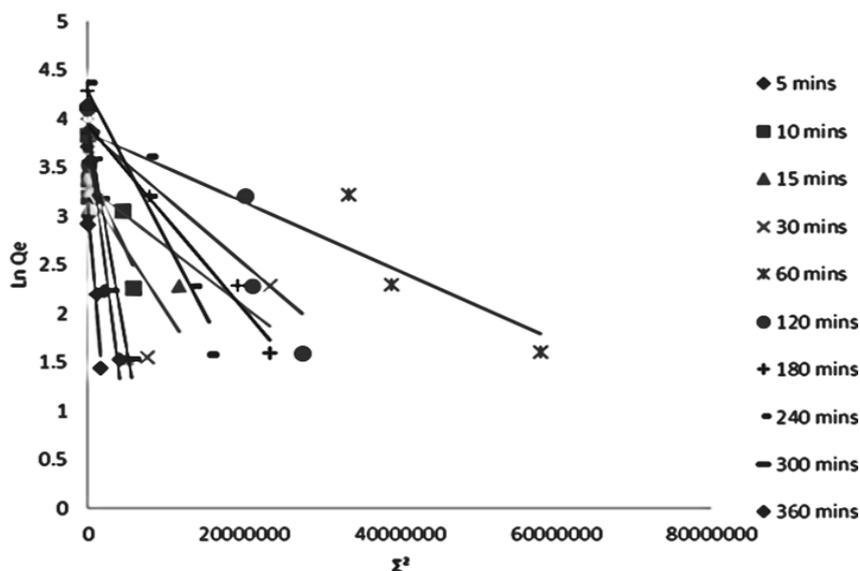


Figure 11 Dubinin-Radushkevich (D-R) isotherm plot for derived chitosan-silica hybrid aerogel in a nickel solution.

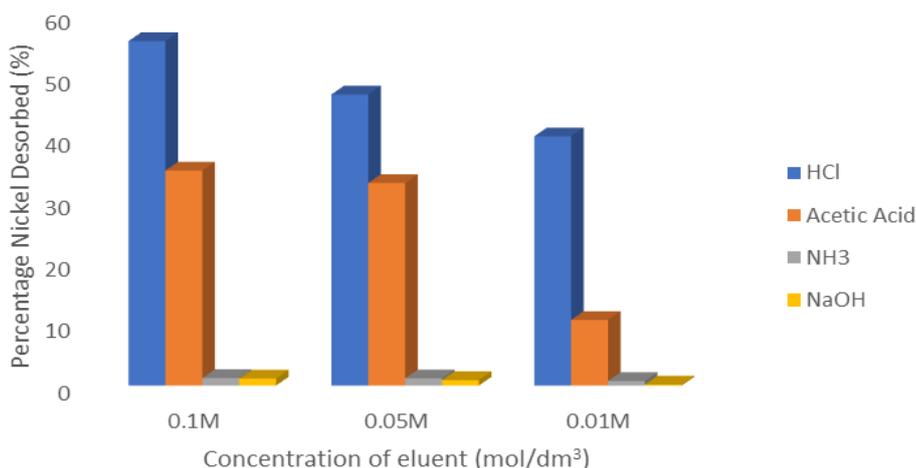


Figure 12 Plot of the percentage of 50 ppm nickel desorbed.

## Conclusions

With an optimum pH of 3 and sorption contact time of 60 min, Nickel ions (99.78 %) were maximally recovered from aqueous solution by (CS)<sub>h</sub>A. When the initial nickel ion concentration was raised from 10 - 200 mg/L, the number of nickel ions removed from the solution ranged from 4.98 - 85.84 mg/g. Langmuir isotherm and Pseudo-second-order kinetics best explain the adsorption process. The separation factor also indicates that the sorption process is favorable under the Langmuir model. Dilute hydrochloric acid (0.1 M) was more effective in recovering bound nickel ions (55.63 %) from the derived aerogel. Hence, nickel ions can be effectively recovered from and by chitosan-silica hybrid aerogel when utilized in contaminated water.

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